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-Intro to stoichiometry Chem Fax Mole Ratios Answers

First we determine the mole ratio (remember unknown species on top), $H_2S : H_2 = 8:8 =$
Number of moles of given species x mole ratio = number of moles required of unknown.
So $3.5 \times 1 = 3.5$ moles of H_2S could be produced. II. What

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would be the mass of this maximum yield? To answer this we need to convert our mole value for H_2S into a mass value.

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ratio of NaClO and $\text{Na}_2\text{S}_2\text{O}_3$ using the method of continuous variations, it was discovered based on the data that this ratio occurred when there was 40 mL of Chem Fax Mole Ratios Answers - restapi205.tasit.com

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moles Cu reacted =

$$0.1/63.5 = 1.574 \times 10^{-3}$$

2. mass of silver = 52.1

- 50.1 = 2.00g. moles

Ag produced = $2/108 =$

8.19×10^{-1} . 3. ratio

Ag:Cu = $8.19 \times 10^{-1} :$

$1.574 \times 10^{-3} = 819 \times$

$10^{-3} : 1.574 \times 10^{-3}$...

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the vinegar used in the

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lab had 5.89 g of acetic acid per 100. mL of ...

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.... ..moles Cu reacted =
 $0.20\text{g} \times (1 \text{ mol} / 63.55\text{g})$
 $= 3.15 \times 10^{-3}$ (2).. mass
Ag produced = $50.67\text{g} -$
 $50.03\text{g} = 0.64\text{g}$
..mole Ag produced =
 $0.64\text{g} \times (1 \text{ mol} / 107.9\text{g})$
 $= 5.93 \times 10^{-3}$ (3)

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data that this ratio
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40 mL of

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