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Bonding- Chemistry by

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Chapter 8 (Bonding:
General Concepts) - Part

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Section 8.2 – The
Nature of Covalent
Bonding In ionic
bonding, atoms transfer
electrons to achieve
noble gas configuration.
In covalent bonding,
atoms share electrons to
achieve noble gas

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configuration. Most atoms share electrons until they have a total of 8 valence electrons (octet rule).

Chapter 8 – Covalent Bonding

Chapter 8: Covalent Bonding Review

Worksheet Section 8.1

Molecular Compounds

1. Classify each of the following as an atom or

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molecule: a. Be b. N₂ c.

CO₂ d. H₂O e. Ne₂.

Which of the following

are diatomic molecules?

a. CO₂ b. N₂ c. O₂ d.

H₂O e. CO₃. What

types of elements tend

to combine to form

molecular compounds?
4.

Covalent bonding

worksheet - Chapter 8

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Bonding and Molecular
Structure 8-6 Ex ample:

PF₃ Ex CO₂ Step 1:

Count Valence

Electrons Count the

total number of valence
electrons in the

molecule or ion. Anions
have extra electrons, so

add 1 e lectron for each

negative charge. Cations
have a deficiency of

electrons, so subtract 1

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electron for each
positive charge.

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Molecular Structure

this covalent bonding
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answers alitaoore can be
taken as competently as
picked to act. covalent
bonding chapter 8
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Covalent Bonding and

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Molecular Structure 8-6

Example: PF₃ Ex CO

2 Step 1: Count Valence

Electrons Count the

total number of valence

electrons in the

molecule or ion.

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Bonding Solutions

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Chapter 8 121 Section

8.1 The Covalent Bond

pages 240–247 Practice

Problems page 244

Draw the Lewis

structure for each

molecule. 1. PH₃ H₂ H₂

H—H H₂ P₂ respectively,

for single, double, and

triple P₂ — — 2. H₂ S H₂ H₂

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H — H S S — 3. HCl

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set (62) A covalent bond

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in which one atom
contributes both
bonding electrons.

Coordinate Covalent
Bond. A bond formed
when two atoms share a
pair of electrons.

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8.6 Lewis Structures
and Formal Charge

- The electron surplus or

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deficit, relative to the
free atom, that is
assigned to an atom in a
Lewis structure. Formal
charges are not “real”
charges. H: orig.
valence e = 1 non-
bonding e = 0 $1/2$
bonding e = 1 formal
charge = 0 O: orig.
valence e = 6 non-
bonding e = 4

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Experimental evidence,
however, indicates ...

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Covalent Bonding and
Molecular Structure 8-6

Example: PF_3 Example CO

2 Step 1: Count Valence

Electrons Count the

total number of valence
electrons in the

molecule or ion. Anions
have extra electrons, so

add 1 electron for each
negative charge.

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conceptual subject and,
in order to explain many
of the concepts, teachers
use models to describe
the microscopic world
and relate it to the
macroscopic properties
of matter. This can lead
to problems, as a

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student's every-day experiences of the world and use of language can contradict the ideas put forward in chemical science. These titles have been designed to help tackle this issue of misconceptions. Part 1 deals with the theory, by including information on some of the key alternative conceptions that have been

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uncovered by research; ideas about a variety of teaching approaches that may prevent students acquiring some common alternative conceptions; and general ideas for assisting students with the development of appropriate scientific conceptions. Part 2 provides strategies for dealing with some of the misconceptions that

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students have, by including ready to use classroom resources including copies of probes that can be used to identify ideas held by students; some specific exercises aimed at challenging some of the alternative ideas; and classroom activities that will help students to construct the chemical concepts required by the

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previous editions of An Introduction to Physical Science, the goal of the new Thirteenth edition is to stimulate students' interest in and gain knowledge of the physical sciences.

Presenting content in such a way that students develop the critical reasoning and problem-solving skills that are needed in an ever-

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changing technological world, the authors emphasize fundamental concepts as they progress through the five divisions of physical sciences: physics, chemistry, astronomy, meteorology, and geology. Ideal for a non-science majors course, topics are treated both descriptively and

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quantitatively, providing instructors the flexibility to emphasize an approach that works best for their students.

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